



Application Note AN-RS-054

# Monitoring phosphate reactions in real time with Raman spectroscopy

## Improving product quality in fertilizer production

Phosphorus and nitrogen are essential inorganic nutrients required for plant growth. While nitrogen is primarily sourced from the atmosphere via fixation processes, phosphorus is obtained by mining phosphate rock. Phosphorus supplements are vital to overcome diminishing crop yields as soil reserves deplete over time [1].

Phosphate fertilizer is produced via a wet chemical process in which phosphate rock is reacted with sulfuric acid to yield phosphoric acid and calcium sulfate dihydrate (gypsum) [2]. Traditional analytical

methods—titration for acid and phosphate quantification, and gravimetric analysis for gypsum—are time consuming, require hazardous reagents, and provide delayed feedback. These limitations hinder real-time process optimization [3].

Raman spectroscopy provides a non-invasive, reagent-free alternative for real-time analysis. It can simultaneously detect phosphate and sulfate species in solution, enabling tighter process control and improved fertilizer product quality.

## CURRENT METHODES FOR REACTION ANALYSIS

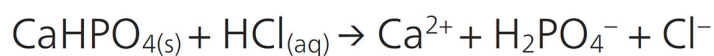
Titration is highly effective for quantifying phosphate and sulfate species to determine reaction completion and final product composition in fertilizers. Gravimetric analysis is used to quantify insoluble byproducts like gypsum ( $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ ). Both techniques require sample collection, hazardous

reagents, and manual processing – all of which introduce time delays and incur costs with every measurement. This Application Note describes how Raman spectroscopy was used as an alternative method to monitor a simplified phosphate fertilizer production process under laboratory conditions.

## METHOD

Raman spectroscopy is well suited to monitor phosphate species due to its high selectivity and nondestructive nature. Furthermore, it can simultaneously detect sulfate ions, enabling real-time analysis of multi-component reaction systems. In this study, dicalcium phosphate ( $\text{CaHPO}_4$ , hereafter referred to as DCP) is used as a model compound for phosphate rock. In a 100 mL beaker, 500 mg of DCP

was dissolved in 10 mL of 0.5 mol/L hydrochloric acid (HCl) (**Equation 1**). To simulate industrial reaction conditions and introduce sulfate ions, 0.25 mL of 1.0 mol/L sulfuric acid ( $\text{H}_2\text{SO}_4$ ) was then added. The use of low-concentration HCl ensures safe handling and facilitates clear identification of the sulfate ( $\text{SO}_4^{2-}$ ) Raman peak.



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**Equation 1**

## METHOD

The resulting acidic mixture was titrated with 1 mol/L sodium hydroxide (NaOH) in 0.25 mL increments while Raman spectra were acquired with 1064 nm Raman (SpecSuite software; 100% laser power, 30 s integration, 1 average). Continuous monitoring of pH using a Metrohm 913 pH meter with Electrode Plus

(6.0262.100) permitted correlation of chemical speciation with spectral changes. Maintaining pH below 4 prevents precipitation of brushite ( $\text{CaHPO}_4 \cdot 2\text{H}_2\text{O}$ ) and other calcium salts from  $\text{HPO}_4^{2-}$  and  $\text{Ca}^{2+}$  (**Equations 2–3**).



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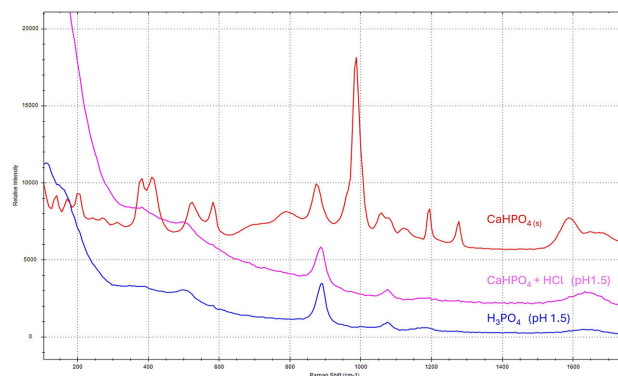
**Equation 2**



### Equation 3

## RESULTS

The Raman spectrum of DCP powder (**Figure 1**) exhibited characteristic peaks consistent with previously reported values [4]. Upon dissolution in 0.5 mol/L HCl, the solution displayed Raman peaks at 889 and 1189  $\text{cm}^{-1}$  corresponding to fully protonated phosphoric acid ( $\text{H}_3\text{PO}_4$ ) and a peak at 1076  $\text{cm}^{-1}$  from the dihydrogen phosphate ion ( $\text{H}_2\text{PO}_4^-$ ). These features closely matched those of a reference  $\text{H}_3\text{PO}_4$  solution, confirming the successful dissolution of DCP and the coexistence of multiple phosphate species under strongly acidic conditions.

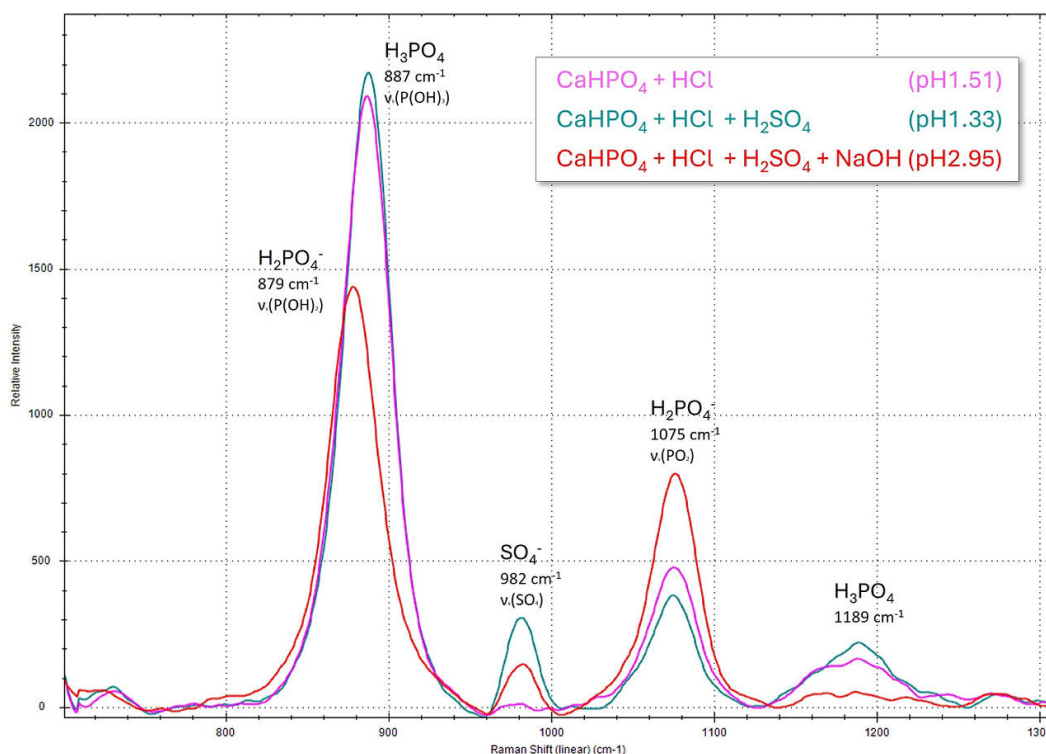


**Figure 1.** Raman spectra of DCP powder (red), DCP dissolved in HCl (pink), and reference phosphoric acid solution (blue).

After adding sulfuric acid, a distinct Raman peak corresponding to  $\text{SO}_4^{2-}$  appeared at 983  $\text{cm}^{-1}$ . In aqueous solution, sulfuric acid dissociates in two steps (**Equation 4**),  $\text{SO}_4^{2-}$  being the predominant species under conditions of low pH and dilution (**Figure 2**).

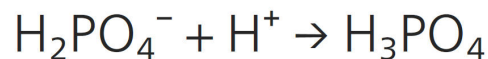


### Equation 4



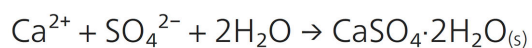
**Figure 2.** Raman spectra of DCP in HCl (pink), after addition of sulfuric acid (green), and after neutralization to pH 2.95 (red).

Phosphate peak intensities changed after H<sub>2</sub>SO<sub>4</sub> addition, indicating protonation of H<sub>2</sub>PO<sub>4</sub><sup>-</sup> to H<sub>3</sub>PO<sub>4</sub> (**Equation 5**) and increased ionic strength. These results demonstrate Raman spectroscopy's ability to detect phosphates and sulfates simultaneously, and its sensitivity to protonation state changes, consistent with [prior findings](#).



**Equation 5**

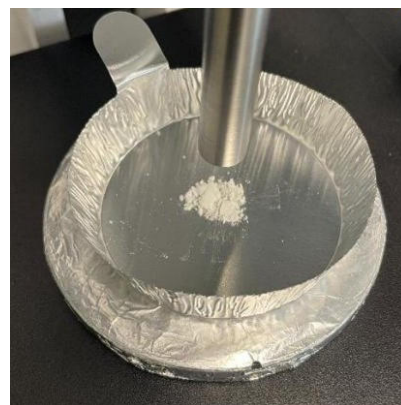
Stepwise addition of 1.0 mol/L NaOH raised the pH, leading to visible precipitation. At pH 2.95, the 889 cm<sup>-1</sup> peak shifted to 879 cm<sup>-1</sup>, the 1189 cm<sup>-1</sup> peak dropped to baseline, and the 1076 cm<sup>-1</sup> peak increased—confirming H<sub>3</sub>PO<sub>4</sub> deprotonation (**Figure 2**). Concurrently, SO<sub>4</sub><sup>2-</sup> intensity fell by 52%, well above the expected 16%, suggesting sulfate removal via precipitation (likely as gypsum, **Equation 6**), a known step in phosphorus fertilizer production to eliminate excess Ca<sup>2+</sup> and SO<sub>4</sub><sup>2-</sup> [4].



**Equation 6**

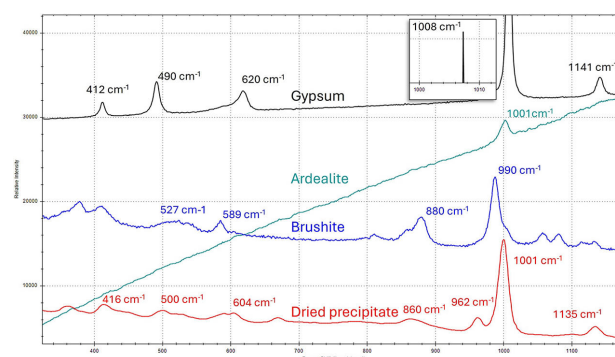


To verify this, the solid precipitate was isolated, washed, dried, and analyzed with Raman spectroscopy (**Figure 3**). The dried sample exhibited a dominant Raman peak at  $1001\text{ cm}^{-1}$ , slightly downshifted from the  $1008\text{ cm}^{-1}$  peak typical of pure gypsum (**Figure 4**). This shift may reflect the co-precipitation of gypsum, brushite, and ardealite ( $\text{Ca}_2(\text{HPO}_4)(\text{SO}_4)\cdot 4\text{H}_2\text{O}$ ) [5]. Ardealite exhibits a  $\text{SO}_4^{2-}$  peak near  $1001\text{ cm}^{-1}$ . These mixed precipitates may be the result of elevated local pH (3.5–4) as NaOH makes contact with the reaction mix.



**Figure 3.** Precipitate recovered from the reaction and mounted on a BAC151B probe holder for Raman analysis.

These results underscore the utility of Raman spectroscopy as a selective tool for monitoring reaction progress and product composition in real time. It has potential applications in both process optimization and quality assurance.



**Figure 4.** Raman spectra of dried precipitate compared with reference spectra of gypsum, brushite, and ardealite. Spectra are vertically offset for clarity.

## CONCLUSION

Raman spectroscopy proved to be an effective tool for monitoring a phosphate fertilizer production process, providing clear insights into phosphate speciation, sulfate concentration, and precipitation dynamics. Raman spectral changes revealed transitions in

phosphate protonation states and the formation of calcium-based precipitates. These capabilities support enhanced process control and product quality in phosphate fertilizer manufacturing.

## REFERENCE

1. Cordell, D.; Drangert, J.-O.; White, S. The Story of Phosphorus: Global Food Security and Food for Thought. *Global Environmental Change* **2009**, *19* (2), 292–305.  
<https://doi.org/10.1016/j.gloenvcha.2008.10.009>.
2. US EPA Office of Air and Radiation (OAR). 8.9 Phosphoric Acid. In *AP 42, Fifth Edition, Volume I Chapter 8: Inorganic Chemical Industry*; AP-42: Compilation of Air Emissions Factors from Stationary Sources; 2020; Vol. 1.  
[https://www.epa.gov/sites/default/files/2020-09/documents/8.9\\_phosphoric\\_acid.pdf](https://www.epa.gov/sites/default/files/2020-09/documents/8.9_phosphoric_acid.pdf)
3. Metrohm AG. *Determination of Total Phosphate in Phosphoric Acid and Phosphate Fertilizers with Thermometric Titration*; AB-314; Metrohm AG.
4. Barua, R.; Daly-Seiler, C. S.; Chenreghanianzabi, Y.; et al. Comparing the Physicochemical Properties of Dicalcium Phosphate Dihydrate ( DCPD ) and Polymeric DCPD ( P-DCPD ) Cement Particles. *J Biomed Mater Res* **2021**, *109* (10), 1644–1655.  
<https://doi.org/10.1002/jbm.b.34822>.
5. Lafuente, Barbara. The Power of Databases: The RRUFF Project. **2015**.  
<https://doi.org/10.1515/9783110417104-003>.

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## CONFIGURATION



### i-Raman NxG 785H

The i-Raman NxG 785H is ideal for routine quality control and process monitoring, especially where speed, stability, and reliability are essential. It offers a balanced combination of performance and efficiency to measure Raman scattering from 100 – 2800  $\text{cm}^{-1}$ . Designed to support high signal throughput, this flexible system is ideal for monitoring chemical and polymer reactions, optimizing processes, and performing content uniformity testing of pharmaceutical tablets. The i-Raman NxG 785H can be easily adapted for see-through measurements even through opaque containers, adding to its versatility.

The i-Raman NxG 785H is the go-to solution for teams looking for reliable Raman analysis in challenging operational environments.

Discover why the i-Raman NxG is the perfect way to gain control of your quality control measurements:

- High-sensitivity spectrometers deliver results in seconds and can detect the faintest Raman signals
- Flexible fiber optic probe compatible with a wide array of accessories, including a vial holder, cuvette holder, immersion probe, and see-through adapter
- Powerful SpecSuite software for easy Raman data collection in addition to quantitative model building, identification with spectral libraries, and routine analysis
- Compact and stackable to save on valuable bench space.